**Stoichiometry / Limiting Reagent Quiz: Honors**

Given the equation below, solve the following questions: (balance = 1 pt)

**\_\_\_ Na3PO4 + \_\_\_ K2O → \_\_\_ K3PO4 + \_\_\_ Na2O**

1. How many grams of sodium oxide can I make from 11.7 grams of sodium phosphate and 20.0 grams of potassium oxide? (10 pt)
2. How many grams of the excess reagent will remain when this reaction has been completed? (5 pt)
3. Before doing the reaction from problem 1, I discovered a big bottle of sodium phosphate in the stock room. Given this near-infinite quantity of sodium phosphate and the 20.0 grams of potassium oxide I already had, how many grams of potassium phosphate can I make? (6 pt)
4. Define the following terms: (2 pt each)

* limiting reagent:
* actual yield:
* stoichiometry: